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In this activity you will explore how mass, temperature, heat energy, and the type of substance are related. a rate of 30 Joules per minute Calorimetry 1 Model 1 - A Pot of Water Before heating 24.5 ° C Temp. 4 qt. saucepan 24.5 ° C Temp. 8 qt. stockpot After heating for 5 minutes at

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Calorimetry – Measurement of Heat Energy 7. Determine the mass of a water sample that is heated from an initial temperature of 25 °C to a final temperature of 100. °C following the addition of 1200 J of heat energy. 8. A 100. g sample of pure lead is heated from 10.0°C to 197.5°C by the addition of 3000. J of heat energy.

Calorimetry – Measurement of Heat Energy

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"The goal of POGIL [Process-orientated guided-inquiry learning] is to engage students in the learning process, helping them to master the material through conceptual understanding (rather than by memorizing and pattern matching), as they work to develop essential learning skills." -- P. v.

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