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Dilutions KEY Molarity and Dilution
Molarity Practice Problems Dilution
Problems - Chemistry Tutorial

Understanding Molarity and Dilution
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Serial Dilution **Dilutions - Part 2 of 4**

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(Serial Dilutions) *Molarity Made Easy: How to Calculate Molarity and Make Solutions* **Dilutions - Part 3 of 4**
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Chemistry Tips How to Calculate
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Stock Solution vs Diluted Solution

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**Molarity, Solutions, Concentrations
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Worksheet Key~~ *Dilution Chemistry:
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Answers # 1 . 11. Molarity Unit Review
2 . 12. Chemistry 11 Calculations
Practice Test # 1 . 13. Chemistry 11
Calculations Practice Test # 2 .
Molarity Worksheet # 1 . 1. 15.8 g of
KCl is dissolved in 225 mL of water.
Calculate the molarity.

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Molarity Worksheet # 1

molarity of the diluted solution be?

$$(0.75 \text{ M})(250 \text{ mL}) = M_2 (295 \text{ mL})$$
$$M_2 = \frac{(0.75 \text{ M})(250 \text{ mL})}{295 \text{ mL}} = 0.64 \text{ M}$$

2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what will the molarity of the diluted solution be? $(0.45 \text{ M})(175 \text{ mL})$

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$$M_1 V_1 = M_2 V_2 \quad (250 \text{ mL}) \quad M_2 = (0.45 \text{ M})(175 \text{ mL}) = 0.32 \text{ M} \quad (250 \text{ mL})$$

~~Dilutions Worksheet W 329 – Everett
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Dilution Problems Worksheet 1. How
do you prepare a 250.-ml of a 2.35 M
HF dilution from a 15.0 M stock

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Answers? 2. If 455-ml of 6.0 M HNO_3 is used to make a 2.5 L dilution, what is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is

~~Molarity Problems Worksheet~~ Mrs

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Molarity WS - HN KEY. Name: Part 1:
Molarity $M = \frac{\text{moles of solute}}{\text{Volume of Solution (L)}}$ Date: Molarity and
Dilutions Practice - = Mols Block: 1. 2.
3. 4. 5. 6. What is the molarity of a
0.30 liter solution containing 0.50
moles of sodium chloride. Calculate

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Answers
the molarity of 0.289 moles of Iron (III) Chloride, FeCl_3 , dissolved in 120 of 1000 FL.

~~Molarity WS - HN KEY~~

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~~Liberty Union High School District /~~

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Overview

$M_1 V_1 = M_2 V_2$ (1.71 M) (25.0 mL)
= M₂ (65.0 mL) M₂ = 0.658 M. M =
mol/L = (25.0/40.0) / (0.325) = 1.92
mol/L. g = (M) (L) (FW) = (0.400) (
(0.225) (119) = 10.7 g. (25.0g) (1
mol/101 g) (1000mL/0.650 mol) = 381
mL. Zn (NO₃)₂ AlCl₃ CuAc₂. 2 mol Ca

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$(0.0500) \times 5.55 \text{ g Ca (OH)}_2 = 0.2775 \text{ g}$

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Dilutions Worksheet - Solutions 1) If I
add 25 mL of water to 125 mL of a
0.15 M NaOH solution, what will the
molarity of the diluted solution be?

$$M_1V_1 = M_2V_2 \quad (0.15 \text{ M})(125 \text{ mL}) = x$$

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(150 mL) $\times = 0.125 \text{ M}$ 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$

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Dilutions Worksheet – Solutions 1) If I

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Answers
have 340 mL of a 0.5 M NaBr solution, what will the concentration be if I add 560 mL more water to it? 0.19 M (the final volume is 900 mL, set up the equation from that) 2) If I dilute 250 mL of 0.10 M lithium acetate solution to a volume of 750 mL, what will the concentration of this solution be?

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Dilutions work, Dilution work answers, Chemistry dilution practice, Dilutions work name key, Solutions work 2 molarity and dilution problems answers.

~~Dilution Worksheets - Kiddy Math~~

This worksheet features 5 molarity

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Answers (M=mol/L) with conversions from grams to moles and milliliters to liters and 7 dilutions problems using $M_1V_1=M_2V_2$. ANSWER KEY INCLUDED! Follow me on Twitter @DenmanChem to see more from my chemistry class.

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This worksheet and quiz will let you practice the following skills: Reading comprehension - ensure that you draw the most important information from the related how to calculate molarity and ...

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Dilution Problems Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO₃ is used to make a 2.5 L dilution, what

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is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is the molarity of the stock solution? 4.

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• **molarity:** the number of moles in a liter (volume), $M = \text{mol/L}$ • equation for dilutions: $M_1V_1 = M_2V_2$, the concentration (or molarity) x volume of your original solution = the new concentration x new volume o In this case, the number of moles stays the same but the volume changes.

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~~Molarity and Serial Dilutions Teacher
Handout~~

This worksheet provides many examples for students to practice calculations involving Molarity & Molality. A complete answer key is provided at the end. This worksheet

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Answers can be used in any Chemistry class,
regardless of the students' ability level.

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