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Chapter 3: Pure Substances Phase Change and Property Diagrams In this chapter we consider the property values and relationships of a pure substance (such as water) which can exist in three phases – solid, liquid and gas.

Chapter 3: Pure Substances – Thermodynamics

The change in internal energy can be found from the first law of thermodynamics: $\Delta U = Q - W = (3.5 \times 10^7 \text{ J}) - (2.1 \times 10^7 \text{ J}) = 0.9 \times 10^7 \text{ J} = 90 \text{ kJ}$. A gas in a cylinder is kept at a constant pressure of $3.5 \times 10^7 \text{ Pa}$ while 300 kJ of heat are added to it, causing the gas to expand from 0.9 m^3 to 1.5 m^3 .

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3-5 3-23 Problem 3-22 is reconsidered. The missing properties of water are to be determined using EES, and the solution is to be repeated for refrigerant-134a, refrigerant-22, and ammonia.

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Chapter 3: THERMODYNAMICS. -Thermodynamics is the study of the relationship between the energy transformation in the system and other physical quantities such as temperature, pressure and volume (P, V, T). -A thermodynamic equation of state is a mathematical relationship of the thermodynamic or state variables, such as pressure, volume and temperature.

Chapter 3: THERMODYNAMICS

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Chapter 3-3 Heat transfer is energy in transition due to a temperature difference. The three modes of heat transfer are conduction, convection, and radiation. Conduction through Plane Walls Conduction heat transfer is a progressive exchange of energy between the molecules of a substance. Fourier's law of heat conduction is $Q = -kA \frac{dT}{dx}$ here Q!

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